



Name:

Class:

Student book questions

3.5 Metal cations and non-metal anions combine to form ionic compounds

Pages 78–79

Check your learning 3.5

Remember and understand

1 Carefully examine the periodic table in Figure 3.9 (page 71).

a Which elements are likely to form positively charged ions?

b Which elements are likely to form negatively charged ions?

c What does this tell you about which elements will combine to form ionic compounds?

2 What kinds of particles are present in ionic compounds?

Apply and analyse

3 How does the group in which an element is found in the periodic table quickly identify one or more of its properties?



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4 What is the maximum number of electrons that can be gained or lost by an atom? Why?

5 Use your knowledge of atomic structure and valence electrons to explain why many ionic compounds are made up of a metal and a non-metal.
